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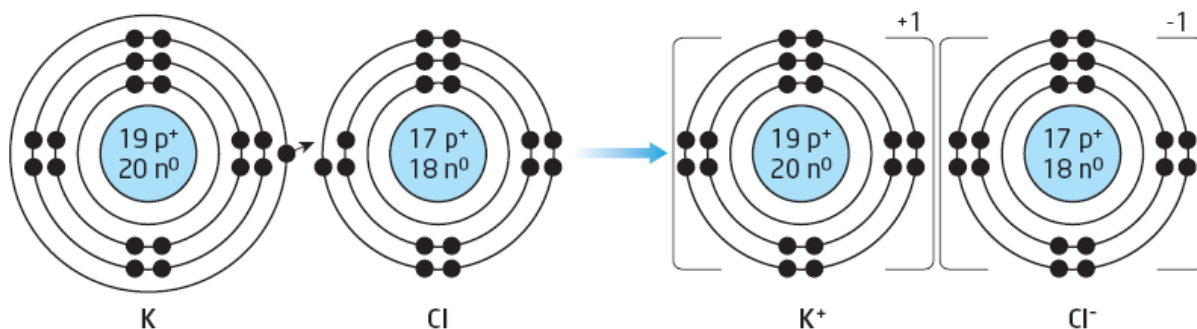
Lesson 6A Ionic Compounds (Nelson p.178-180)

Learning Goals:

- A. Identify simple ionic compounds, using the periodic table, and write the formulae (C3.8)

How Ionic Compounds Form

- It is made up of a _____ and a _____.
- **Metals** _____ **electrons** to non-metals, to form _____.
- **Non-metals** _____ **electrons** from metals, to form _____.
- The bond that forms between them is called an _____.



Example: Potassium (K) and Sulfur (S) combine to form an ionic compound.

Use the Chart below:

Step 1: Draw a Bohr-Rutherford Diagram for Potassium in the metal side of the chart.

Step 2: Draw a Bohr-Rutherford Diagram for Potassium in the metal side of the chart.

Step 3: Determine how many Potassium atoms are required to make sulfur stable.

Use the chart below to complete the questions:

1. Calcium (Ca) combine with chlorine (Cl)
2. Lithium (Li) combine with oxygen (O)
3. Beryllium (Be) combine with sulfur (S)
4. Sodium (Na) combine with fluorine (F)
5. Magnesium (Mg) combine with oxygen (O)

Metal	Metal	Non-metal	Non-metal	Chemical Formula
Potassium (K)	Potassium (K)	Sulfur (S)		K_2S

Metal	Metal	Non-metal	Non-metal	Chemical Formula
1. Calcium (Ca)		Chlorine (Cl)	Chlorine (Cl)	
2. Lithium (Li)	Lithium (Li)	Oxygen (O)		
3. Beryllium (Be)		Sulfur (S)		
4. Sodium (Na)		Fluorine (F)		
5. Magnesium (Mg)		Oxygen (O)		