

Name: _____

Date: _____

Worksheet 3C Electron Arrangements & Reactivity (Nelson p.166)

1. Use the **periodic table** to complete the chart:

Element	Metal / Non-metal	Group	How many outer electron(s)	Lose / Gain how many e to be stable
Oxygen				
Sodium				
Chlorine				
Aluminum				
Magnesium				
Bromine				
Nitrogen				
Argon				
Sulfur				
Potassium				

2. **Name** and write the **symbol** for these elements:

Description	Name	Symbol
a) the noble gas with the smallest atomic number		
b) the non-metal in the period 3 with 6 outer electrons		
c) the metal in the period 2 with 1 outer electron		
d) the Group 14 element that has 2 orbits		
e) the element with atomic number 33		
f) the element with mass number 80		
g) the element with atomic mass 137.33		
h) the stable element with 10 protons		
i) the halogen in period 2		
j) the alkali metal that has 12 neutrons		
k) the alkaline earth metal that has 20 electrons		
l) the element in group 13, period 4		